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a hydronium ( $\text{H}_3\text{O}^+$ ) and a hydroxide ( $\text{OH}^-$ ) ion  
 $2 \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  The ionization constant ( $K_w$ )  
of water is:  $K_w = [\text{H}_3\text{O}^+] [\text{OH}^-] = 1.0 \times 10^{-14}$   
M Acidic, Basic, & Neutral IF  $[\text{H}_3\text{O}^+] >$

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CHAPTER 15 REVIEW Acid-Base Titration and pH  
SECTION 2 SHORT ANSWER Answer the following  
questions in the space provided. 1. Below is  
a pH curve from an acid-base titration. On it  
are labeled three points: X, Y, and Z. a.  
Which point represents the equivalence point?  
b. At which point is there excess acid in the  
system? c.

### ~~15 Acid Base Titration and pH~~

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in this set (10) Self-ionization of water.  
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pH. the negative of the common logarithm of  
the hydronium ion concentration. pOH.

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1 Modern Chemistry Chapter 12-Solutions 2

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