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the subshell before
pairing electrons. 2.

packets/quanta of

electromagnetic energy 3.

the SI unit of frequency

4. An atomic orbital can
hold no more than two

electrons. 5. the amount
of energy required to

move an electron from its
present energy level to

the next higher one 6. the

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this set (82) Rutherford's
model-also called nuclear
model-did not explain
how the electrons were
arranged Chapter 5:
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- 1) Electrons circulate around the nucleus
- 2) Electrons can jump from one level to the next but can NOT exist between levels (ladder)
- 3) Energy levels are NOT evenly spaced (escalator)
- 4) Energy lost or gained is NOT always the same
- 5) The higher the energy level, the easier for the electrons to leave (magnet)

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~~Answers~~
the formula $2n^2$. the
maximum number of
electrons that can occupy
an energy level. opposite
spin. in order to occupy
the same orbital 2
electron must have. filled
energy sub levels. stable
electron configurations
are likely to contain. all

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the orbitals contain one electron, with spins parallel.

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states that single electrons with the same spin must occupy each equal-energy orbital before additional electrons with opposite spins can occupy the same orbitals.

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Pauli exclusion principle. states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite spins. photoelectric effect.

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Wave-Particle

Duality/Electromagnetic

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