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The measurement has two significant figures.
The measurement has four significant figures.
Decimal point is present, so these zeros are significant. The rules are summarized in the following flowchart:

Nonzero digits 3 4 0 0 g
of sulfur pp Nonzero

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digits 4 0 0 0. mL of
oxygen pp

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General Plan for
Solving Stoichiometry
Problems Convert using
the mole ratio A, given
in the balanced
chemical equation. Mass
of substance A Amount
in mol of substance A
Amount in mol of
substance B Convert
using the molar mass of

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A. 1 Mass of substance
B 4 2 3 Convert using
the molar mass of B. B
SAMPLE PROBLEM 1

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Stoichiometry

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SAMPLE PROBLEM 1

What is the molarity of a solution that is made by diluting 50.00 mL of a 4.74 M ...

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1. Determine the

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percentage composition
of each of the following
compounds: a. sodium
oxalate, $\text{Na}_2\text{C}_2\text{O}_4$
ans: 34.31% Na,
17.93% C, 47.76% O b.
ethanol, $\text{C}_2\text{H}_5\text{OH}$
ans: 52.13% C, 13.15%
H, 34.72% O c.
aluminum oxide, Al_2O_3
ans: 52.92% Al,
47.08% O d. potassium
sulfate, K_2SO_4 ans:
44.87% K, 18.40% S,

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36.72% O2.

Solving Answer

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CHAPTER 6 ...

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CHAPTER 14

Concentration of

Solutions There are

three principal ways to

express solution

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concentration in chemistry—percentage by mass, molarity, and molality. The following table compares these three ways of stating solution concentration. Examining the method of preparation of the three types may

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CHAPTER 14...

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Neither pressure nor temperature is constant; therefore, the combined gas law must be used to calculate the volume of CO₂ at the final temperature and pressure. multiply by the volume ratio, CO₂ use the combined gas law to

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convert from the initial
temperature and
pressure to the final
temperature and
pressure C 2 H 4

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CHAPTER 18

Titration

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CHAPTER Empirical
Formulas

ADDITIONAL
PROBLEMS

PRACTICE. Name

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3. Determine the
empirical formula for
compounds that have

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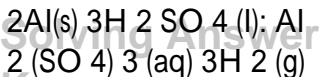
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3. Aluminum will react

with sulfuric acid in the

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following reaction.



a. How many moles of H_2SO_4 will react ans: 27 mol H_2SO_4 with 18 mol Al? b. How many moles of each product ans: 27 mol H_2 will be produced? 29 mol $\text{Al}_2(\text{SO}_4)_3$.

SAMPLE PROBLEM 1

Ammonia is made

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industrially by reacting

...
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CHAPTER 18

Titration Chemists
have many methods for
determining the

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quantity of a substance present in a solution or other mixture. One common method is titration, in which a solution of known concentration reacts with a sample containing the substance of unknown quantity. There are two main requirements

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