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Chapter 11 Review

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between the nu... It applies only to gases measured under the same temperature a....

chapter 11 test chemistry molecular composition Flashcards ...

334 CHAPTER 11 FIGURE 11-1 At the same temperature and pressure, balloons of equal volume have equal numbers of molecules, regardless of which gas they contain. Hydrogen molecule 1 mol H₂ at STP = 22.4 L
Oxygen molecule 1 mol O₂ at STP = 22.4 L
Carbon dioxide molecule 1 mol CO₂ at STP = 22.4 L

CHAPTER 11 Molecular Composition of Gases

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Section 11.2 Answers
The atmosphere contains 78% nitrogen gas, 21% oxygen gas, and 1% other gases by volume. Atmospheric pressure decreases higher up in the atmosphere.

Chapter 11 Test Molecular

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Chapter 11 - Molecular Composition of Gases 11-1 Volume-Mass

Relationships of Gases I. Measuring and Comparing the Volumes of Reacting Gases A. Observations of Gay-Lussac
1. 2 liters H_2 + 1 liter O_2 \rightarrow 2 liters H_2O vapor
2. 2 volumes H_2 + 1 volume O_2 \rightarrow 2 volumes H_2O vapor
3. 1 volume H_2 + 1 volume Cl_2 \rightarrow 2 volumes HCl
4.

Chapter 11 - Molecular Composition of Gases

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CHAPTER 11 REVIEW Molecular

Composition of Gases MIXED REVIEW

SHORT ANSWER Answer the

following questions in the space

provided 1 c The average speed of a gas molecule is most directly related to the (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) Chapter 11 Review Gases Mixed Answer Key

Chapter 11 Review Gases Mixed

Answers | calendar.pridesource

Chapter 11 Review Molecular

Composition Of Gases Mixed Answers

Answer. Molar mass = 151.1 g/mol

Potassium: $(39.10 / 101.1) \times 100 =$

38.67% Nitrogen: $(14.01 / 101.1) \times$

100 = 13.86% Oxygen: $(48.00 /$

101.1) $\times 100 = 12.58\%$.

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Google Sites The atmosphere contains

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and 1% other gases by volume.

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Atmospheric pressure decreases higher up in the atmosphere. Chapter 11 Review Molecular Composition Of ...

Molecular Composition Of Gases 11 3
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CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided 1 The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol (c) multiplied by 224 L (b) divided by the mass of 1 mol (d) divided by 224 L 2 For the expression $V = nR/P T$

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